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2)At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases boyle's law 1)The relationship between the pressure and volume of a gas at constant temperture; when volume increase, pressure decreases.,

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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as $P_{CO_2} = 0.285$ torr and $P_{N_2} = 593.525$ torr. What is the partial pressure of oxygen?

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SECTION 4 Date CHAPTER 11 REVIEW Gases Class SHORT ANSWER Answer the following questions in the space provided. Increasing rate of effusion. List the following gases in order (Assume all gases are at the same temperature and pressure.) (c) HCl (d) 2. Explain your reasoning for the order of gases you chose in item 1 above. Refer to the kinetic-

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A 500.0 mL gas sample at STP is compressed to a volume of 300.0 mL and the temperature is increased to 35.0°C. What is the new pressure of the gas in pascals? 8. A sample of gas occupies 1000. mL at standard pressure. What volume will the gas occupy at a pressure of 600. mm Hg if the temperature remains constant? 96 GASES MODERN CHEMISTRY

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