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gas was 12.0 L. What was the initial volume of the gas? Notice: as the temperature doubled, what happened to the volume?! ! $V_1 = 6.0 \text{ L}$ c) The volume of a gas was originally 2.5 L; its pressure was 104 kPa and its temperature was 270K. The volume of the gas expanded to 5.3 L and its pressure decreased to 95 kPa. What is the temperature of the ...

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5. Some oxygen gas has a volume of 41.0 L under a pressure of 245 kPa and a temperature of 279 K. What is the mass of the gas? 138.6 g
6. 17.5 mL of oxygen gas were collected at room temperature (22°C) and 100.2 kPa of atmospheric pressure.
a) How many moles of oxygen gas were produced? $7.15 \times 10^{-4} \text{ mol}$

Information: Gas Pressure

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